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CHM134 GENERAL CHEMISTRY I
VSEPR worksheet #1

Complete the following table

Example	# valence electrons	Lewis Structure	Electronic Geometry	Molecular Geometry	a. o. hybrid.	Draw (angles)
CBr_4	32		Td	Td	sp^3	
ClF_3	28		Tbp	T shaped	sp^3d	
BCl_3	24		Trig. Pl.	Trig Pl	sp^2	
KrF_2	22		Tbp	linear	sp^3d	
SF_4	34		Tbp	sec-saw	sp^3d	
PCl_6^{-1}	48		Oh	Oh	sp^3d^2	

IONIC BONDING 3 Name: _____
Ionic bonding occurs between a metal and a non-metal. In the chart below show the transfer of electrons with an arrow.

$\text{K} \cdot + \cdot \text{Br} \rightarrow \text{K}^+ \text{Br}^-$ $\text{Mg} \cdot + \cdot \text{Cl} \cdot \rightarrow \text{Mg}^{2+} \text{Cl}_2^{2-}$

IONIC BONDING 2 Name: _____
(Additional Practice)
Ionic bonding occurs between a metal and a non-metal. In the chart below show the transfer of electrons with an arrow.

Fe^{+2}

IONIC BONDING 1 Name: _____
Ionic bonding occurs between a metal and a non-metal. In the chart below show the transfer of electrons with an arrow.

$\text{Na} \cdot + \cdot \text{Cl} \cdot \rightarrow \text{Na}^+ \text{Cl}^-$ $\text{Mg} \cdot + \cdot \text{Cl} \cdot \rightarrow \text{Mg}^{2+} \text{Cl}_2^{2-}$

1. Fe^{+2}	1. $\text{Na} +$	1. $\text{K} + \text{O}$	$\text{K} \cdot + \cdot \text{O} \cdot \rightarrow \text{K}^+ \text{O}^{2-}$	K_2O
2. Fe^{+3}	2. $\text{Ag} +$	2. $\text{Mg} + \text{O}$	$\text{Mg} \cdot + \cdot \text{O} \cdot \rightarrow \text{Mg}^{2+} \text{O}^{2-}$	MgO
3. $\text{Pb} +$	3. $\text{Ca} +$	3. $\text{Ca} + \text{S}$	$\text{Ca} \cdot + \cdot \text{S} \cdot \rightarrow \text{Ca}^{2+} \text{S}^{2-}$	CaS
4. Pb^{+2}	4. $\text{Zn} +$	4. $\text{Na} + \text{Br}$	$\text{Na} \cdot + \cdot \text{Br} \cdot \rightarrow \text{Na}^+ \text{Br}^-$	NaBr
5. Cu^+	5. $\text{Ba} +$	5. $\text{H} + \text{Cl}$	$\text{H} \cdot + \cdot \text{Cl} \cdot \rightarrow \text{H}^+ \text{Cl}^-$	HCl
6. Cu^+	6. $\text{Al} +$	6. $\text{Al} + \text{Cl}$	$3 \text{Al} \cdot + 3 \cdot \text{Cl} \cdot \rightarrow 3 \text{Al}^+ 3 \text{Cl}^-$	AlCl_3
7. Cr^+	7. $\text{Al} +$	7. $\text{Ni} + \text{O}$	$2 \text{Ni} \cdot + 3 \cdot \text{O} \cdot \rightarrow 2 \text{Ni}^{2+} 3 \text{O}^{2-}$	Ni_2O_3

DRAW A LEWIS DOT STRUCTURE

STEP 1 **NO₃⁻**
Nitrogen: 5e
Oxygen: 6e (3e)
Ion's charge: 1e
24 electrons
COUNT ALL VALENCE ELECTRONS

STEP 2 **0 electrons**
DRAW THE LEAST ELECTRONEGATIVE ATOM IN THE CENTER WITH THE OTHER ATOMS SPACED EVENLY AROUND IT

STEP 3 **6 electrons**
DRAW SINGLE BONDS TO THE CENTRAL ATOM

STEP 4 **24 electrons**
COMPLETE OCTETS WITH REMAINING ELECTRONS FOR THE OUTER ATOMS FIRST AND ADD REMAINING ELECTRONS (IF ANY) TO THE CENTRAL ATOM
NOTE 1: FOR HYDROGEN USE THE DUET RULE

STEP 5 **24 electrons**
COMPLETE THE CENTER ATOM'S OCTET BY MAKING COVALENT BONDS USING LONE PAIR ELECTRONS FROM AN OUTER ATOM OF YOUR CHOICE

STEP 6 - DRAWING RESONANCE STRUCTURES & DETERMINING FORMAL CHARGE
formal charge of an atom = [# valence electrons] - [# dots] - [# lines]

DEPENDENT ON WHICH ELECTRONS YOU MOVE IN STEP 5
A VARIETY OF RESONANCE STRUCTURES CAN BE DETERMINED
NOTE 2: THE "CORRECT" STRUCTURE CONSIDERS ALL RESONANCES AND CAN ONLY BE DETERMINED EXPERIMENTALLY
PERIOD 1 ELEMENTS FOLLOW THE DUET RULE - PERIOD 2 ELEMENTS FOLLOW THE OCTET RULE - PERIOD 3 ELEMENTS CAN HAVE EXPANDED OCTETS

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